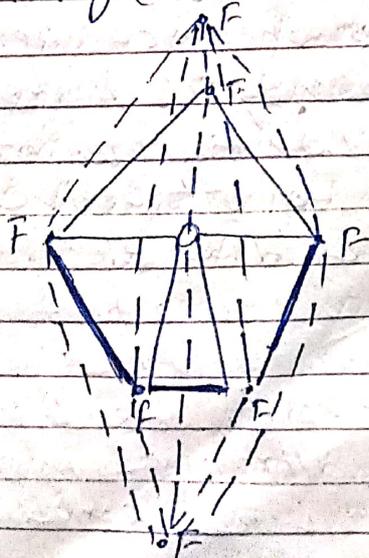


(e) Six bonding electron pair: \rightarrow These are 6 bonding electron pair in molecule the molecule will be octahedral in shape and bond angle will be 90° , ex - SF_6 . central atom S has 6 V.E., these are used in bond formation with 6 F-atom. Since all the 6 bonded pairs are equivalent & no. V.E. is left unused. The mole. has regular geometry of an octahedral noted at (2d).

(f) Seven bonding electron pair: \rightarrow If there are 7 bonding electron pair, the molecule will be pentagonal bipyramidal, ex - IF_7 . The str. of molecule can be better if we visualise. two pyramids placed in such a way that their base, touch one another, the base of pyramid is pentagonal in fig (3).



(1 F7) fig (3).